



The Mole Concept
Chemistry – Leaving Cert
Quick Notes

The Mole Concept

The relative molecular mass of a substance is the mass of one molecule of that substance compared with one twelfth of the mass of the carbon-12 isotope. It is calculated by adding together the atomic masses of the elements present. One mole of a substance is the amount of that substance which contains 6×10^{23} particles of that substance. 6×10^{23} is known as Avogadro's constant. A mole is the relative atomic mass or the relative molecular mass expressed in grams. The number of moles in a substance can be calculated by dividing the mass in grams by the mass of one mole.